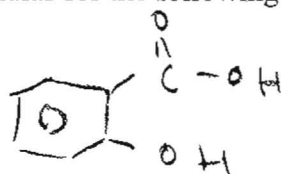


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Student ID# \_\_\_\_\_

1. Show valid structural formulas for the following compounds (6 pts)

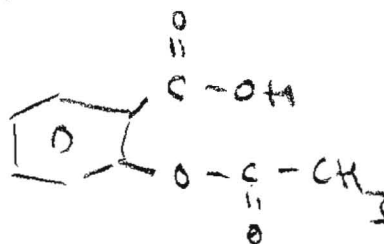
a. salicylic acid



b. ethanol



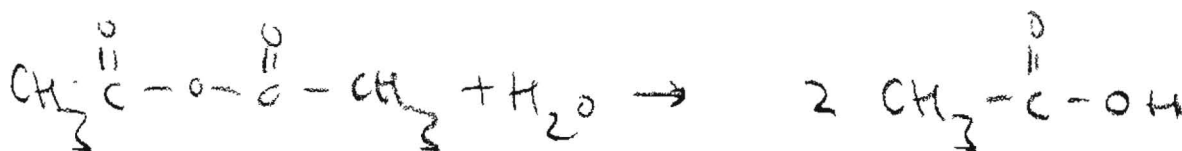
c. aspirin



2. If you take 11.0 g of aspirin, how many moles of aspirin does that correspond to?  
 (Atomic weights: C = 12.011, H = 1.008, O = 16.00) 4 pts

$$\text{MW } \text{C}_9\text{H}_8\text{O}_4 = 180.16 \quad \frac{11.0}{180.2} = 0.061 \text{ moles}$$

3. Show a balanced reaction equation for the reaction of acetic anhydride with water to acetic acid! Use structural formulas. (3 pts)



4. How many moles of sulfuric acid did your reaction consume for each mole of aspirin prepared? (2 pts).

0 (NONE!)